Chemistry 115 Name

Dr. Cary Willard

Exam 2b March 16, 2011

Multiple Choice (30 points)

Page 5 (20 points)

Page 6 (25 points)

Page 7 (16 points)

Page 8 (10 points)

Total (101 points)

All work must be shown to receive credit. Give all answers to the correct number of significant figures

Avogadros number = 6.022 x 1023 /mol

Grossmont College

Periodic Table

|  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
| IA |  |  |  |  |  |  |  |  |  |  | |  |  |  |  |  | VIIA | NOBLE GASES |
| 1  **H**  1.008 | IIA |  |  |  |  |  |  |  |  |  | |  | IIIA | IVA | VA | VIA | 1  **H**  1.008 | 2  **He**  4.002 |
| 3  **Li**  6.941 | 4  **Be**  9.012 |  |  |  |  |  |  |  |  |  | |  | 5  **B**  10.81 | 6  **C**  12.01 | 7  **N**  14.01 | 8  **O**  16.00 | 9  **F**  19.00 | 10  **Ne**  20.18 |
| 11  **Na**  23.00 | 12  **Mg**  24.30 | IIIB | IVB | VB | VIB | VIIB | VIII VIII VIII | | | | IB | IIB | 13  **Al**  27.00 | 14  **Si**  28.09 | 15  **P**  30.97 | 16  **S**  32.06 | 17  **Cl**  35.45 | 18  **Ar**  39.95 |
| 19  **K**  39.10 | 20  **Ca**  40.08 | 21  **Sc**  44.96 | 22  **Ti**  47.90 | 23  **V**  50.94 | 24  **Cr**  52.00 | 25  **Mn**  54.94 | 26  **Fe**  55.85 | 27  **Co**  58.93 | 28  **Ni**  58.70 | | 29  **Cu**  63.55 | 30  **Zn**  65.38 | 31  **Ga**  69.72 | 32  **Ge**  72.59 | 33  **As**  74.92 | 34  **Se**  78.96 | 35  **Br**  79.90 | 36  **Kr**  83.80 |
| 37  **Rb**  85.47 | 38  **Sr**  87.62 | 39  **Y**  88.91 | 40  **Zr**  91.22 | 41  **Nb**  92.91 | 42  **Mo**  95.94 | 43  **Tc**  (99) | 44  **Ru**  101.1 | 45  **Rh**  102.9 | 46  **Pd**  106.4 | 47  **Ag**  107.9 | | 48  **Cd**  112.4 | 49  **In**  114.8 | 50  **Sn**  118.7 | 51  **Sb**  121.8 | 52  **Te**  127.6 | 53  **I**  126.9 | 54  **Xe**  131.3 |
| 55  **Cs**  132.9 | 56  **Ba**  137.3 | 57  **La**  138.9 | 72  **Hf**  178.5 | 73  **Ta**  180.9 | 74  **W**  183.9 | 75  **Re**  186.2 | 76  **Os**  190.2 | 77  **Ir**  192.2 | 78  **Pt**  195.1 | 79  **Au**  197.0 | | 80  **Hg**  200.6 | 81  **Tl**  204.4 | 82  **Pb**  207.2 | 83  **Bi**  209.0 | 84  **Po**  (209) | 85  **At**  (210) | 86  **Rn**  (222) |
| 87  **Fr**  (223) | 88  **Ra**  226.0 | 89  **Ac**  227.0 | 104  **Rf**  (261) | 105  **Db**  (262) | 106  **Sg**  (263) | 107  **Bh**  (262) | 108  **Hs**  (265) | 109  **Mt**  (266) | 110  **??**  (269) |  | |  |  |  |  |  |  |  |

|  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
| 58  **Ce**  140.1 | 59  **Pr**  140.9 | 60  **Nd**  144.2 | 61  **Pm**  (147) | 62  **Sm**  150.4 | 63  **Eu**  152.0 | 64  **Gd**  157.3 | 65  **Tb**  158.9 | 66  **Dy**  162.5 | 67  **Ho**  164.9 | 68  **Er**  167.3 | 69  **Tm**  168.9 | 70  **Yb**  173.0 | 71  **Lu**  175.0 |
| 90  **Th**  232.0 | 91  **Pa**  231.0 | 92  **U**  238.0 | 93  **Np**  (237) | 94  **Pu**  (244) | 95  **Am**  (243) | 96  **Cm**  (247) | 97  **Bk**  (247) | 98  **Cf**  (251) | 99  **Es**  (252) | 100  **Fm**  (257) | 101  **Md**  (258) | 102  **No**  (259) | 103  **Lr**  (260) |

Lanthanide series

Actinide series

Part I – Multiple Choice (30 points)

1. When ultraviolet energy is absorbed by an atom, an electron
   1. moves from a lower to a higher energy level.
   2. moves from a higher to a lower energy level.
   3. is ejected from an atom completely.
   4. falls into the nucleus of an atom.
   5. is taken on by an atom.
2. Any *p* sublevel can hold up to \_\_\_\_\_\_\_\_ electrons.
   1. 2
   2. 6
   3. 8
   4. 18
   5. 10
3. Sulfur is a \_\_\_\_\_\_\_\_ block element.
   1. *d*
   2. *f*
   3. *s*
   4. *p*
   5. *g*
4. The atomic radius of bromine is larger than the atomic radius of *\_\_\_\_\_\_\_\_.*
   1. uranium
   2. potassium
   3. chlorine
   4. iodine
   5. xenon
5. Neon has \_\_\_\_\_\_\_\_ valence electrons.
   1. 2
   2. 4
   3. 8
   4. 6
   5. 10
6. In ionic compounds, \_\_\_\_\_\_\_\_ lose their valence electrons to form positively charged \_\_\_\_\_\_\_\_.
   1. metals, cations
   2. metals, anions
   3. nonmetals, cations
   4. metals, polyatomic ions
   5. nonmetals, anions
7. Which one of the following elements forms two or more ions with different ionic charges?
   1. K
   2. F
   3. Ca
   4. O
   5. Fe
8. Which of the following polyatomic ions has a 3- ionic charge?
   1. phosphate
   2. hydroxide
   3. nitrate
   4. sulfate
   5. bicarbonate
9. A sodium ion is an example of a(n) \_\_\_\_\_\_\_\_.
   1. anion
   2. cation
   3. polyatomic ion
   4. nonmetal ion
   5. neutral atom
10. According to the IUPAC nomenclature system, the types of compound that use prefixes in their names are \_\_\_\_\_\_\_\_.
    1. ionic compounds
    2. covalent compounds
    3. ionic compounds involving transition metals
    4. polyatomic ions
    5. compounds that contain polyatomic ions
11. Which of the following properties is most characteristic of an organic molecule?
    1. contain sodium
    2. soluble in water
    3. has a high melting point.
    4. has C-H bonds
    5. has ionic bonds
12. One mol of particles of any substance contains how many particles?
    1. 106
    2. 3 × 10-10
    3. 3 × 1010
    4. 6.022 × 1023
    5. 6.022 × 10-23
13. A chemical equation is balanced when
    1. the total number of molecules is the same in reactants and products.
    2. the total number of ions is the same in reactants and products.
    3. the sum of the coefficients of the reactants is equal to the sum of the coefficients of the products.
    4. the number of atoms of each element is the same in reactants and products.
    5. the charge on each atom is the same in reactants and products.
14. The reaction of carbon with oxygen to produce carbon monoxide is an example of which class of reaction?

2C (*s*) + O2 (*g*) → 2CO (*g*)

* 1. single replacement
  2. combination
  3. double replacement
  4. catalytic
  5. endothermic

1. In any balanced chemical equation, the number of each type of atom on both sides of the equation is \_\_\_\_\_\_\_\_.
   1. the same
   2. doubled
   3. decreased by one
   4. increased by one
   5. dependent on the temperature

Part 2 – Problems and Short Answer (75 points)

1. (3 points) How do chemists explain the lines of color that are seen in the atomic spectra of the elements?

The lines of color are caused by the energy released as electrons drop from higher energy levels down to lower energy levels. This energy is released as light of a specific wavelength (color).

1. (3 points) Write the complete electron configuration of sulfur.

1s2 2s2 2p6 3s2 3p4

1. (5 points) Write the shorthand electron configuration of rhodium (Rh)

[Kr] 5s2 4d7

Write the shorthand configuration for a Rh+2 ion.

[Kr] 4d7

1. (5 points) Describe how ionization energies change as you move across the periodic table to the right and explain the reason for this trend.

As you move across the table to the right, the effective charge felt by the outermost electrons increases thereby pulling them closer to the nucleus and thus the electrons are held more tightly.

1. (4 points) Circle the element with the larger atomic radius

Chlorine or Iodine ?

Bromine or Calcium ?

1. (6 points) Name the following compounds

|  |  |  |  |
| --- | --- | --- | --- |
|  | Cation name | Anion name | Compound name |
| Ba(NO3)2 | Barium ion | Nitrate ion | Barium nitrate |
| Fe(OH)3 | Iron(III) ion | Hydroxide ion | Iron(III) hydroxide |
| Ag2S | Silver ion | Sulfide ion | Silver sulfide |
| P2S3 |  |  | Diphosphorus trisulfide |

1. (6 points) Give the correct formula for the following compounds

|  |  |  |  |
| --- | --- | --- | --- |
|  | Cation formula | Anion formula | Compound formula |
| Ammonium chloride | NH4+ | Cl-1 | NH4Cl |
| Lead(IV) sulfate | Pb+4 | SO4-2 | Pb(SO4)2 |
| Aluminum oxide | Al+3 | O-2 | Al2O3 |
| Tetrasulfur octabromide |  |  | S4Br8 |

1. (4 points) Calculate the number of moles of lead in a 153 g sample of lead (Pb).
2. (4 points) Calculate the number of atoms of lead in 3.24 moles of lead.
3. (5 points) Determine the empirical formula of ethyl butyrate, the principle component of pineapple oil. It is composed of 62.04% C, 10.41% H and 27.55% O.
4. (4 points) A compound has an empirical formula of C3H2N and a molar mass of 312.3 g/mol. What is the molecular formula of the compound?

Molar mass of C3H2N = 3(12) + 2(1) + 14 = 52 g/mol

There are 312/52 or 6 units of this in the compound

Molecular formula = C18H12N6

1. (3 points) Calculate the molar mass of methyl salicylate(C8H8O3), the ester responsible for the aroma of wintergreen
2. (4 points) Calculate the mass of 0.813 moles of methyl salicylate.
3. (5 points) Calculate the number of atoms of carbon in 5.00 g of methyl salicylate.

or

1. (4 points) Balance the following equations
   1. 4 HCl + O2 🡪 2 H2O + 2 Cl2
   2. 2 C8H14 + 23 O2 🡪 16 CO2 + 14 H2O
2. (2 points) Name the following hydrocarbon.

 . 3-methyl heptane

1. (4 points) Match each of the following molecules to the correct functional group.
   1.  alcohol (1)
   2.  alkyne (4)
   3.  carboxylic acid (8)
   4.  ketone (11)

Alcohol

Aldehyde

Alkene

Alkyne

Amide

Amine

Aromatic

Carboxylic acid

Ester

Ether

Ketone